

Automatic Translation?,,?,,?,,?,,?,,?

Figure 1. (A) Schematic of the principle of insertion and disinsertion in the electrode materials. (B). Structure of lithium ion batteries. Reproduced with permission from Nishi (2001).

Table 1. Historical account of the major breakthroughs in lithium ion batteries discovery.

Figure 2. (A) Battery production in Japan for the year 2013. Secondary batteries (rechargeable) represent 39% while primary batteries (including lithium metal, zinc silver oxide and Zn-MnO₂) represent 61%. Adapted from Julien et al. (2016). (B) The worldwide rechargeable battery market, in volume, MWh, 1995–2015, reproduced with permission from Pillot⁴. (C) Forecasts for automobile rechargeable battery market, M\$ for x-EV 2000–2025, reproduced with permission from Pillot³.

Figure 4. (A) Schematic design of lithium ion battery pack with aluminum foam between heat spreader and the battery. Adapted from Saw et al. (2017). (B) Remote battery monitoring centers linking BMS to big data facilities. Adapted from Xiong et al. (2018).

Figure 5. (a–c) “Lithium metal deposition at high current density (≈50 mA cm−²) propagates into densely pressed polycrystalline β-Li₃PS₄ from a brass tip electrode. (d) Viewed in transmission optical microscopy, the lithium metal network shows a branching pattern. (e) The starting as-pressed surface of the β-Li₃PS₄ polycrystal exhibits sub-micrometer (≈200 nm length scale) cracks or pores.” Both figure and caption are reproduced with permission from Porz et al. (2017).

Figure 6. (A) Schematic illustration of the potential vs. Li/Li+ and the corresponding capacity density of active present anode materials for lithium ion batteries. Reproduced with permission from Goriparti et al. (2014). (B) Safety tests on a full cell using nail penetration (a) Carbon/LMO and (b) MSNP-LTO/LMO cells. (c,d) Overcharge tests displaying T_{surface} and V of Carbon/LMO (c) and MSNP-LTO/LMO (d). Reproduced with permission from Amine et al. (2010).

Table 2. Safety and capacity of the common anodic materials used in LiB. Reproduced with minor changes with permission from Goriparti et al. (2014).

Figure 7. (A) Thermal stability of the common lithium ion batteries cathode materials⁶. (B) Accelerated Rate Calorimetry data of a 18,650 full cell displaying normalized self-heating rate. Improving cathode stability results in a higher thermal runaway temperature (increased stability) and a slow heating rate. Printed with permission from Roth and Doughty (2006).

Figure 8. Energy densities and energy efficiencies of the most promising electrode materials Reproduced with permission from Schmuck et al. (2018).

Table 3. Thermodynamic characteristics of common solvents in which inorganic salts are dissolved to form organic electrolytes.

Figure 10. (A) Structure of hydrofluoro-electrolyte, adapted from Nagasubramanian and Fenton (2013). (B) Dual electrolyte concept scheme adapted from Wang et al. (2015). (C) Lithium ionic conductivities of most common solid-state electrolytes. The highlighted zone represents the thermal stability domain of liquid electrolytes. Printed with permission from Park et al. (2018).

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