Fuel cell characteristics

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Figure 1 shows a simplified diagram of a typical fuel cell, which has three basic elements:

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A fuel cell can be defined as an electrochemical cell that generates electrical energy from fuel via an electrochemical reaction.

A fuel cell is an electrochemical cell that generates electrical energy from fuel via an electrochemical reaction. It offers high efficiency and zero emissions.

A fuel cell is different from the conventional methods of energy generation because, in a fuel cell, chemical energy is directly converted into electrical energy without intermediate conversion into mechanical power.

A fuel cell is preferred over conventional methods of energy generation because, in a fuel cell, zero combustion takes place. Thus, carbon dioxide is not produced.

Lithium potassium carbonate salt is used as an electrolyte in molten carbonate fuel cells.

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A fuel cell is a device that generates electricity through an electrochemical reaction, not combustion. In a fuel cell, hydrogen and oxygen are combined to generate electricity, heat, and water. Fuel cells are used today in a range of applications, from providing power to homes and businesses, keeping critical facilities like hospitals, grocery stores, and data centers up and running, and moving a variety of vehicles including cars, buses, trucks, forklifts, trains, and more.

Fuel cell systems are a clean, efficient, reliable, and quiet source of power. Fuel cells do not need to be periodically recharged like batteries, but instead continue to produce electricity as long as a fuel source is provided.

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