

# Is gibbs free energy extensive

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Gibbs free energy, also known as the Gibbs function, Gibbs energy, or free enthalpy, is a quantity that is used to measure the maximum amount of work done in a thermodynamic system when the temperature and pressure are kept constant. Gibbs free energy is denoted by the symbol  $\Delta G$ . Its value is usually expressed in Joules or Kilojoules. Gibbs free energy can be defined as the maximum amount of work that can be extracted from a closed system.

This property was determined by American scientist Josiah Willard Gibbs in the year 1876 when he was conducting experiments to predict the behaviour of systems when combined together or whether a process could occur simultaneously and spontaneously. Gibbs free energy was also previously known as "available energy." It can be visualised as the amount of useful energy present in a thermodynamic system that can be utilised to perform some work.

Gibbs free energy is equal to the enthalpy of the system minus the product of the temperature and entropy. The equation is given as:

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