

# What is lithium chloride used for

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Sources: Lithium chloride - Wikipedia, Lithium Chloride - Properties, Structure, Preparation and Uses - Vedantu, LiCl - Lithium Chloride Structure, Molecular Mass, Properties and Uses.

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Until the 1990s, the metal and lithium market was dominated by US production from mineral deposits.

However, at the beginning of the 21st century most of the production was derived from non-US sources, with Australia, Chile and Portugal being the most important suppliers in the world.

As a curiosity to stake out that, although Bolivia has half of the world's lithium deposits, it is not a major lithium producer.

The most important commercial form is lithium carbonate,  $\text{Li}_2\text{CO}_3$ , Produced from minerals or brines by several different processes.

The addition of hydrochloric acid (HCl) produces lithium chloride. This method produces the hydrated compound (water bound to the molecule). The hydrated form can be dried to the anhydrous form by heating with hydrogen chloride gas.

Lithium chloride is a white solid with no particular aroma and salty taste (National Center for Biotechnology Information, 2005). Their appearance is shown in figure 2.

Lithium chloride are highly hygroscopic crystals with octahedral geometry similar to that of sodium chloride. Its crystalline structure is presented in figure 3 (Winter, S.F.).

Its molecular weight is 42.39 g / mol, its density is 2.068 g / mL and the melting and boiling points are 605 °C and 1360 °C respectively.

The compound is very soluble in water, alcohol, ether, pyridine and nitrobenzene (Royal Society of Chemistry, 2015).

Lithium chloride has weak oxidizing or reducing power. However, redox reactions can still occur. Lithium chloride solutions are generally neither strongly acidic nor strongly basic.

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